

ERRATUM

A. ARCORA, F. P. BALLISTRERI and G. A. TOMASELLI: Reactions of 2-thiophenesulphonyl chloride with anion and neutral nucleophiles. Solvent effects on nucleophilic reactivity correlations. *Tetrahedron* 14(16), 2545-2556 (1978).

The author regrets miscalculations in Tables 1, 3, 4, 5, 6.

A revised Table 1 is printed below, new values in Tables 3-6 can be calculated from the revised Table 1.

Table 1

Second order rate constants k_2 for the reactions of 2-thiophenesulphonyl chloride with nucleophiles Nu in various solvents at 25°C.

$$k_2 \cdot 10^3(a) \text{ (l mol}^{-1} \text{ sec}^{-1}\text{)}$$

Solvents	Nu= NO ₂ ⁻	N ₃ ⁻	S ₂ O ₃ ²⁻	F ⁻	AcO ⁻	C ₅ H ₅ N	C ₆ H ₅ NH ₂	Im.	H ₂ O
1 Water ^(b)	25.9	163	1015	20.3	2.55	220	709 ⁽¹⁰⁾	125	0.008598
2 Heavy Water	27.8	-	-	-	3.08	167	589	128	-
3 Methanol ^(c)	1.56	0.916	12.6	0.827	0.529	1.76	9.41	11.0	0.00412
4 Ethanol ^(c)	1.84	1.89	-	0.715	0.181	1.87	5.45	9.42	0.00153
5 Propan-1-ol ^(c)	1.40	1.35	-	-	0.136	0.877	3.12	-	0.00151
6 2-Methoxyethanol ^(c)	10.2	0.267	-	-	0.695	2.69	3.77	-	0.00122
7 Acetonitrile ^(c)	118	0.853	-	-	0.257	12.7	3.33	92.6 ^(d)	0.00165
8 Acetonitrile ^(e)	131	-	-	-	1.59	13.6	5.86	225	-
9 Acetone ^(c)	173	1.02	-	-	0.366	3.75	1.00	61.7	0.000133
10 Dioxane ^(c)	21.4	-	-	-	2.99	1.68	0.753	32.3	0.00290
11 Dioxane ^(f)	20.2	0.140	-	3.06	1.94	30.01	27.1	108	0.000704

(a) Values of k_2 are reproducible to $\pm 6\%$

(b) In this solvent the values of k_2 for p-anisidine and SCN⁻ are 2.317 and 0.000426 respectively.

(c) Mixtures with 10%(v/v) of water.

(d) In pure acetonitrile the value of k_2 is 0.195.

(e) Mixture containing D₂O in equal amount to (c).

(f) Mixture with 40%(v/v) of water. In this solvent the values of k_2 for morpholine and piperidine are 3.00 and 11.00 respectively.